

Lesson 2 Definitions

- Atom – small particles that make up everything around us (matter).
- Element – If a substance is made up of only one type of atom. Eg. Zinc, Oxygen, Copper.
- Compound – a substance that contains two or more different elements that are chemically combined. Eg. Copper Sulphate, Water.
- Mixture – a substance that contains two or more different substances that are not chemically combined.
- Proton – a positive charged object held within the nucleus of an atom. Neutron – an uncharged object held within the nucleus of an atom.
- Electron – a negatively charged object that is free to move around the nucleus.
- Atomic Number – the number of protons in an atom.
- Mass Number – the number of protons and neutrons in an atom.
- Atomic Number – the average Mass Number of all the different isotopes of an element.
- Isotope – when two atoms have the same number of protons but a different number of neutrons (therefore a different mass number).
- Ion – an element or a compound, which takes on a charge due a variation between the number of protons and electrons, it possesses.
- Ionic Charge – the charge of an atom. Eg. Oxygen contains 8 protons and 10 electrons, therefore 8 positive charges and 10 negative charges, i.e. 2- charge
- Cations - are ions, which have a positive charge (due to losing electrons)
- Anions - are ions, which have a negative charge (due to gaining electrons)
- Polyatomic ions - are groups of atoms that have a charge and act as a unit. Eg. SO_4^{2-} contains 1 Sulphur and 4 Oxygen atoms and carries a 2- charge.
- Diatomic elements - are two atoms of the same element bonded together. Eg. H_2 , I_2 , Br_2 , Cl_2 , O_2 , N_2 , Hg_2

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